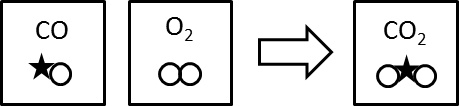
**Chemistry: *Balancing Equations Activity***  Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

*Directions*: Listen to the instructions of your teacher and follow the example below to help you   
 understand how to complete this activity.

\_\_\_\_ CO + \_\_\_\_ O2 🡪 \_\_\_\_CO2



Use the pieces provided to balance each of the following equations. Then fill in the table.

|  |  |  |  |
| --- | --- | --- | --- |
|  | **How many atoms of**  **each type participate**  **in the reaction?** | | |
| **Equation** | **C** | **H** | **O** |
| \_\_\_\_ CH4 + \_\_\_\_ O2 🡪 \_\_\_\_ H2O + \_\_\_\_CO2 |  |  |  |
| \_\_\_\_ H2 + \_\_\_\_ O2 🡪 \_\_\_\_ H2O |  |  |  |
| \_\_\_\_ H2 + \_\_\_\_ O2 🡪 \_\_\_\_ H2O2 |  |  |  |
| \_\_\_\_ CH4 + \_\_\_\_ O2 🡪 \_\_\_\_ CH3OH |  |  |  |
| \_\_\_\_ C6H12O6 + \_\_\_\_ O2 🡪 \_\_\_\_ H2O + \_\_\_\_CO2 |  |  |  |
| \_\_\_\_ C3H8 + \_\_\_\_ O2 🡪 \_\_\_\_ H2O + \_\_\_\_CO2 |  |  |  |
| \_\_\_\_ CO2 + \_\_\_\_ H2 🡪 \_\_\_\_ H2CO + \_\_\_\_O2 |  |  |  |

|  |  |
| --- | --- |
| My rules for balancing C,H,O compounds:  Step 1) balance carbon first  Step 2) balance hydrogen next  Step 3) balance oxygen last | How to balance a chemical equation:  1) polyatomic ions first  2) even / odd (make all odd 🡪 even)  3) single elements last  4) 5/2 (O2) = 5 oxygen and then double equation |