

# NOMENCLATURE REVIEW

## Molecular Compounds, Ionic Compounds, & Acids

NAME THE FOLLOWING COMPOUNDS:

1.  $\text{BaSO}_3$
2.  $(\text{NH}_4)_3\text{PO}_4$
3.  $\text{PBr}_5$
4.  $\text{MgSO}_4$
5.  $\text{CaO}$
6.  $\text{H}_3\text{PO}_4$
7.  $\text{Na}_2\text{Cr}_2\text{O}_7$
8.  $\text{MgO}$
9.  $\text{SO}_3$
10.  $\text{Cu}(\text{NO}_3)_2$
11.  $\text{HI}$
12.  $\text{N}_2\text{O}$
13.  $\text{MnO}$
14.  $\text{AgNO}_3$
15.  $\text{As}_2\text{O}_5$
16.  $\text{Fe}_2\text{O}_3$
17.  $\text{HClO}$
18.  $\text{N}_2\text{O}_3$
19.  $\text{HF}$
20.  $\text{H}_2\text{C}_2\text{O}_4$
21.  $\text{NaHCO}_3$
22.  $\text{SiBr}_4$
23.  $\text{CuCl}_2$
24.  $\text{HNO}_2$
25.  $\text{SnO}_2$
26.  $\text{BaCrO}_4$

WRITE FORMULAS FOR THE FOLLOWING COMPOUNDS:

27. hydrobromic acid
28. chromium(III) carbonate
29. magnesium sulfide
30. iodine trichloride
31. lithium hydride
32. ammonium hydroxide
33. calcium chloride
34. hydroselenic acid
35. iron(II) nitride
36. aluminum hydroxide
37. tin(II) fluoride
38. sulfur tetrachloride
39. mercury(II) iodide
40. diphosphorus pentoxide
41. sulfurous acid
42. lead(II) nitrate
43. dihydrogen monoxide
44. sodium oxalate
45. perchloric acid
46. chlorous acid
47. silicon dioxide
48. carbonic acid
49. sodium chlorate
50. xenon hexafluoride
51. nickel nitrate
52. potassium perchlorate

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### ANSWER KEY

- |                            |                                       |
|----------------------------|---------------------------------------|
| 1. barium sulfite          | 27. HBr                               |
| 2. ammonium phosphate      | 28. $\text{Cr}_2(\text{CO}_3)_3$      |
| 3. phosphorus pentabromide | 29. MgS                               |
| 4. magnesium sulfate       | 30. $\text{ICl}_3$                    |
| 5. calcium oxide           | 31. LiH                               |
| 6. phosphoric acid         | 32. $\text{NH}_4\text{OH}$            |
| 7. sodium dichromate       | 33. $\text{CaCl}_2$                   |
| 8. magnesium oxide         | 34. $\text{H}_2\text{Se}$             |
| 9. sulfur trioxide         | 35. $\text{Fe}_3\text{N}_2$           |
| 10. copper(II) nitrate     | 36. $\text{Al}(\text{OH})_3$          |
| 11. hydroiodic acid        | 37. $\text{SnF}_2$                    |
| 12. dinitrogen monoxide    | 38. $\text{SCl}_4$                    |
| 13. manganese(II) oxide    | 39. $\text{HgI}_2$                    |
| 14. silver nitrate         | 40. $\text{P}_2\text{O}_5$            |
| 15. diarsenic pentoxide    | 41. $\text{H}_2\text{SO}_3$           |
| 16. iron(III) oxide        | 42. $\text{Pb}(\text{NO}_3)_2$        |
| 17. hypochlorous acid      | 43. $\text{H}_2\text{O}$              |
| 18. dinitrogen trioxide    | 44. $\text{Na}_2\text{C}_2\text{O}_4$ |
| 19. hydrofluoric acid      | 45. $\text{HClO}_4$                   |
| 20. oxalic acid            | 46. $\text{HClO}_2$                   |
| 21. sodium bicarbonate     | 47. $\text{SiO}_2$                    |
| 22. silicon tetrabromide   | 48. $\text{H}_2\text{CO}_3$           |
| 23. copper(II) chloride    | 49. $\text{NaClO}_3$                  |
| 24. nitrous acid           | 50. $\text{XeF}_6$                    |
| 25. tin(IV) oxide          | 51. $\text{Ni}(\text{NO}_3)_2$        |
| 26. barium chromate        | 52. $\text{KClO}_4$                   |