		Name:		
Cla a sa	into Doutint November to	Hour: _	Date:	
Cnem	nistry: Partial Neutralization			
<u>Direction</u>	ons: Solve the following problems. Show your work and include u	nits on you	ur answers.	
1.	7.35 L of 0.411 M HI were mixed with 10.26 L of 0.289 M NaOH.	. Find the բ	pH of the resulting solution.	
2.	$68.1~\text{mL}$ of $2.75~\text{M}$ HNO $_3$ were mixed with $61.0~\text{mL}$ of $3.11~\text{M}$ KO	H. Find the	e pH of the resulting solution	
3.	13.8 L of 4.28 M sulfuric acid were mixed with 40.1 L of 2.96 M or resulting solution.	cesium hyd	droxide. Find the pH of the	
4.	27.2 L of 3.81 M hydrobromic acid were mixed with 72.8 L of sochaving a pH of 3.11. What was the concentration of the base?	dium hydro	oxide to produce a solution	

ANSWERS: 1. 2.50 2. 12.28 3. 12.02 4. 1.42 M